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McLaughlin, Kimberly / Thermochemistry

162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: $101.3 \text{ J } 0.18 \text{ L atm } 1 \text{ L atm} = - \cdot \times = \cdot w -18 \text{ J } 6.17$ An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.) If the system absorbs heat, q must be

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+127 J. The change in energy (internal

CHAPTER 6 THERMOCHEMISTRY

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