

## Chapter 19 Electrochemistry Answers

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552 CHAPTER 19: ELECTROCHEMISTRY 19.11 Half-reaction  $E^\circ(\text{V})$   $\text{Mg}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Mg}(\text{s}) -2.37$   $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s}) +0.34$  The overall equation is:  $\text{Mg}(\text{s}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{Cu}(\text{s})$   $E^\circ = 0.34 \text{ V} - (-2.37 \text{ V}) = 2.71 \text{ V}$  19.12 Strategy: At first, it may not be clear how to assign the electrodes in the galvanic cell. From Table 19.1 of

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Chapter 19 Electrochemistry Chang 4 19.4 SPONTANEITY OF REDOX REACTIONS Electrical work is needed to move a charge through a conductor: Electrical energy = charge x potential difference Units: Joules = Coulombs x Volts F is Faraday's constant; the electrical charge contained in 1 mol of e<sup>-</sup> is equal to 96500 C. 1 F = mole<sup>-</sup> Coulombs 1 96500 = V mole<sup>-</sup>

### CHAPTER 19 ELECTROCHEMISTRY

Read Chapter 18: Entropy, Free Energy, and Equilibrium & Read Chapter 19: Electrochemistry Answer the following problems in the space provided. For problems involving an equation, carry out the following steps: 1. Write the equation. 2. Substitute numbers and units. 3. Show the final answer with units. There is no credit without showing work.

### **AP Chemistry Chapter 18 & 19: Thermodynamics ...**

Chapter 19 Electrochemistry Math Summary. Relating Standard Cell Potential to Standard Half Cell Potentials  $E^{\circ}_{\text{cell}} = E^{\circ}_{\text{oxidation}} + E^{\circ}_{\text{reduction}}$  (standard conditions assume 1.0 M concentrations) Relating Half Cell Potentials when Written in Opposite Directions  $E^{\circ}_{\text{ox}} = -E^{\circ}_{\text{red}}$  for half reactions written in opposite directions.

### **Chapter 19 Electrochemistry Math Summary**

48 26 Chem 162-2008 Final Exam + Answers Chapter 18 - Electrochemistry Electrolysis One difference between an electrolytic cell and a galvanic (voltaic) cell is A. electrons flow from the anode to the cathode in a galvanic cell and in the opposite direction in a voltaic cell. B. electrolytic cells always require inert electrodes. C. the galvanic cell reaction is spontaneous and the ...

### **Chapter 18 Electrochemistry Free energy and cell potential ...**

AP Chemistry Review Questions - Electrochemistry. ...  $2.19 \times 10^4$  sec ?  $3.90 \times 10^3$  sec ?  $1.68 \times 10^2$  sec; If a constant current of 8.00 amperes is passed through a cell containing  $\text{Zn}^{2+}$  for 2.00 hours, how many grams of zinc will plate out onto the cathode? ? 39.0 g ?

### **AP Chemistry Review Questions - Electrochemistry**

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Chapter 19 electrochemistry worksheet 4. When the following oxidation-reduction reaction in acidic solution is balanced, what is the lowest whole-number coefficient for  $\text{H}^+$ , and on which side of the balanced equation should it appear?

### **Front Door - Valencia College**

Chapter 19: Electrochemistry Overview of the Chapter review oxidation-reduction chemistry basics galvanic cells spontaneous chemical reaction generates a voltage set-up of galvanic cell & identification of: anode (and half-reaction) cathode (and half-reaction) net cell reaction cell potential ( $E$  or  $E^{\circ}$ ) Chemical thermodynamics (thermochemistry ...