

## Chapter 9 Stoichiometry Review Answers Section 2

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### Chapter 9 Stoichiometry Review Answers

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$   
4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of  $C_3H_4$ ? 2 mol O 2:1 mol H 2O c. What is the mole ratio of O 2 to H

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N<sub>2</sub> are mixed with 120 mol of H

### **Chapter 9 Stoichiometry Review Answers**

Chapter 9 - Stoichiometry 9-1 Introduction to Stoichiometry Composition Stoichiometry - deals with mass relationships of elements in compounds Reaction Stoichiometry - Involves mass relationships between reactants and products in a chemical reaction I. Reaction Stoichiometry Problems A. Four problem Types, One Common Solution

### **Chapter 9 - Stoichiometry**

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### Chapter 9 Review Stoichiometry Answers

Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below:  $2H_2O_2(aq) \rightarrow 2H_2O(l) + O_2(g)$  This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

### Chapter 9: Standard Review Worksheet

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### **Chapter 9: Stoichiometry Review Flashcards | Quizlet**

Stoichiometry b. Theoretically, how many moles of  $\text{NH}_3$  will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the ...

### **Date. FCHAPJ REV[EW.**

Play this game to review Chemistry. Avogadro's number is: Q. Using the pictured equation, how many grams of zinc chloride are produced from 7.89 moles of zinc?

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