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the quantum mechanical model where electrons do not follow specific paths, but are given probabilities of being found in a particular area-- orbitals, clouds; recognized wave-particle duality Orbital the area in which an electron has a 90% probability of being found

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CHEMISTRY NOTES - Chapter 13 Electrons in Atoms Goals : To gain an understanding of : 1. Atoms and their structure. 2. The development of the atomic theory. 3. The quantum mechanical model of the atom. 4. Electron configurations NOTES: The different historical models are described as

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Bridwell, Kim / Chapter 13 - Electrons in Atoms

Chapter 13 - Electrons in Atoms. Chapter 13: 1 - 20, 23 - 25, 27, 31, 32, 34 - 38, 41, 45, 47, 48, 52. Section 13.1 -Models of the Atom. Section Review 13.1. 1. List in chronological order, a major contribution of each of these scientists to the understanding of the atom: proposed that all elements are composed of atoms.

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Chemistry chapter 13. STUDY. PLAY. Amplitude. The height of a wave from the origin to the crest. ... The arrangement of electrons around the nucleus of an atom in its ground state. Energy level. A region around the nucleus of an atom where an electron is likely to be moving. Frequency.

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Chemistry Chapter 13. STUDY. PLAY. quantum. the amount of energy required to move an electron from its present energy level to the next higher one. atomic orbitals. the regions around the nucleus within which the electrons have the highest probability of being found. energy level.

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As given in Table 13.3, the structure of the complex is as shown below: In this complex, atom has 8 electrons outside its noble gas core. The CE present in the complexes is may be CO, CS, CSe, or CTe. The bonding behavior of all these ligands is same.

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33. An atomic orbital may contain 2 electrons at most, and the electrons must have different spins.
35. The highest-energy electrons for all of the elements in groups 1 (1A) and 2 (2A) in the periodic table are in s orbitals. 37. The last electrons to be added to an orbital diagram for the atoms of the transition metal elements go into d orbitals.

Chapter 11 Modern Atomic Theory - An Introduction to Chemistry

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