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3+. (i) The energy level diagram showing the electrons in the five 3d orbitals of a chromium atom is represented below. Draw the completed diagram showing the d orbitals in $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$ after splitting. [1] E (ii) State and explain what happens to the splitting of the d orbitals if the ligand is changed from H_2O to NH_3 .

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1. e i concentration of copper $0.0080 \ll \text{mol dm}^{-3} \gg$ mass of copper in 250.0 cm^3 « $0.0080 \text{ mol dm}^{-3} \times 0.2500 \text{ dm}^3 \gg 63.55 \text{ g mol}^{-1} \gg 0.127 \ll \text{g} \gg$ OR mass of brass in 1 dm^3 «4 0.200 g » 0.800 g AND $[\text{Cu}^{2+}] = \ll 0.0080 \text{ mol dm}^{-3} \gg 63.55 \text{ g mol}^{-1} \gg 0.5084 \text{ g dm}^{-3}$ «% copper in this sample of brass 0.127 0.200

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